Physics

DATA SHEET

Charge on electron, q_e	$-1.602 \times 10^{-19} \mathrm{C}$
Mass of electron, $m_{\rm e}$	$9.109 \times 10^{-31} \text{ kg}$
Mass of neutron, m_n	$1.675 \times 10^{-27} \text{ kg}$
Mass of proton, $m_{\rm p}$	$1.673 \times 10^{-27} \text{ kg}$
Speed of sound in air	340 m s^{-1}
Earth's gravitational acceleration, g	9.8 m s^{-2}
Speed of light, c	$3.00 \times 10^8 \mathrm{ms^{-1}}$
Electric permittivity constant, ε_0	$8.854 \times 10^{-12} \text{ A}^2 \text{ s}^4 \text{ kg}^{-1} \text{ m}^{-3}$
Magnetic permeability constant, μ_0	$4\pi \times 10^{-7} \mathrm{N}\mathrm{A}^{-2}$
Universal gravitational constant, G	$6.67 \times 10^{-11} \mathrm{N}\mathrm{m}^2\mathrm{kg}^{-2}$
Mass of Earth, $M_{\rm E}$	$6.0 \times 10^{24} \mathrm{kg}$
Radius of Earth, $r_{\rm E}$	$6.371 \times 10^{6} \text{ m}$
Planck constant, <i>h</i>	$6.626 \times 10^{-34} \mathrm{Js}$
Rydberg constant, R (hydrogen)	$1.097 \times 10^7 \mathrm{m}^{-1}$
Atomic mass unit, <i>u</i>	$1.661 \times 10^{-27} \text{ kg}$ 931.5 MeV/ c^2
1 eV	$1.602 \times 10^{-19} \mathrm{J}$
Density of water, ρ	$1.00 \times 10^3 \mathrm{kg} \mathrm{m}^{-3}$
Specific heat capacity of water	$4.18 \times 10^3 \mathrm{Jkg^{-1}K^{-1}}$
Wien's displacement constant, b	$2.898 \times 10^{-3} \mathrm{mK}$

FORMULAE SHEET

Motion, forces and gravity

$s = ut + \frac{1}{2}at^2$	v = u + at
$v^2 = u^2 + 2as$	$\vec{F}_{net} = m\vec{a}$ $W = F_{ }s = Fs\cos\theta$
$\Delta U = mg\Delta h$ $P = \frac{\Delta E}{\Delta t}$	$K = \frac{1}{2}mv^2$
$\sum \frac{1}{2}mv_{\text{before}}^2 = \sum \frac{1}{2}mv_{\text{after}}^2$	$P = F_{\parallel}v = Fv\cos\theta$ $\sum m\vec{v} = -\sum m\vec{v}$
$\Delta \vec{p} = \vec{F}_{\text{net}} \Delta t$	$\sum m \vec{v}_{\text{before}} = \sum m \vec{v}_{\text{after}}$ $a_{\text{c}} = \frac{v^2}{r}$
$\omega = \frac{\Delta \theta}{t}$	$F_{\rm c} = \frac{mv^2}{r}$
$\tau = r_{\perp}F = rF\sin\theta$ $2\pi r$	$F = \frac{GMm}{r^2}$
$v = \frac{2\pi r}{T}$ $U = -\frac{GMm}{r}$	$\frac{r^3}{T^2} = \frac{GM}{4\pi^2}$
$U = -\frac{1}{r}$	

Waves and thermodynamics

$v = f\lambda$	$f_{\text{beat}} = \left f_2 - f_1 \right $
$f = \frac{1}{T}$	$f' = f \frac{\left(v_{\text{wave}} + v_{\text{observer}}\right)}{\left(v_{\text{wave}} - v_{\text{source}}\right)}$
$d\sin\theta = m\lambda$	$n_1 \sin \theta_1 = n_2 \sin \theta_2$
$n_{\rm x} = \frac{c}{v_{\rm x}}$ $I = I_{\rm max} \cos^2 \theta$	$\sin\theta_{\rm c} = \frac{n_2}{n_1}$
$Q = mc\Delta T$	$I_1 r_1^2 = I_2 r_2^2$
	$\frac{Q}{t} = \frac{kA\Delta T}{d}$

FORMULAE SHEET (continued)

Electricity and mag	netism
$E = \frac{V}{d}$	$\vec{F} = q\vec{E}$
$V = \frac{\Delta U}{q}$	$F = \frac{1}{4\pi\varepsilon_0} \frac{q_1 q_2}{r^2}$
W = qV	$I = \frac{q}{t}$
W = qEd	V = IR
$B = \frac{\mu_0 I}{2\pi r}$	P = VI
$B = \frac{\mu_0 NI}{L}$	$F = qv_{\perp}B = qvB\sin\theta$ $F = lI_{\perp}B = lIB\sin\theta$
$\Phi = B_{ }A = BA\cos\theta$	$\frac{F}{I} = \frac{\mu_0}{2\pi} \frac{I_1 I_2}{r}$
$\varepsilon = -N \frac{\Delta \Phi}{\Delta t}$	$l = 2\pi r$ $\tau = nIA_{\perp}B = nIAB\sin\theta$
$\frac{V_{\rm p}}{V_{\rm s}} = \frac{N_{\rm p}}{N_{\rm s}}$	$V_{\rm p}I_{\rm p} = V_{\rm s}I_{\rm s}$

Quantum, special relati	vity and nuclear
$\lambda = \frac{h}{mv}$ $K_{\text{max}} = hf - \phi$	$t = \frac{t_0}{\sqrt{\left(1 - \frac{v^2}{c^2}\right)}}$
$\lambda_{\max} = \frac{b}{T}$ $E = mc^2$	$l = l_0 \sqrt{\left(1 - \frac{v^2}{c^2}\right)}$
E = hf	$p_{\rm v} = \frac{m_0 v}{\sqrt{\left(1 - \frac{v^2}{c^2}\right)}}$
$\frac{1}{\lambda} = R\left(\frac{1}{n_{\rm f}^2} - \frac{1}{n_{\rm i}^2}\right)$	$V(-c^{-})$ $N_{\rm t} = N_0 e^{-\lambda t}$
	$\lambda = \frac{\ln 2}{\frac{t_1}{2}}$

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1.008 Hydrogen		г					L	KEY			L						4.003 Helium
∞ .	4 g					Atom	nic Number Symbol	79 114				νœ	٥٢	ΓZ	∞C	6н	10 Ne
6.941 Lithium	9.012 Beryllium					Standard Atomic Weight Name	mic Weight Name	197.0 Gold				10.81 Boron	12.01 Carbon	14.01 Nitrogen	16.00 ^{Oxygen}	19.00 Fluorine	20.18 Neon
11 Na	12 Mo	1									1	13 A1	21 21 22	, 7 1 d	S 16	51	18 Ar
22.99 Sodium	24.31 Magnesium											26.98 Aluminium	28.09 Silicon	30.97 Phosphorus	32.07 Sulfur	35.45 ^{Chlorine}	39.95 Argon
19 K	Ca Ca	21 Sc	22 Ti	23 V	24 Cr	25 Mn	26 Fe	27 C0	28 Ni	29 Cu	30 Zn	31 Ga	Ge Ge	33 As	34 Se	35 Br	36 Kr
39.10 Potassium	40.08 Calcium	44.96 Scandium	47.87 Titanium	50.94 Vanadium	52.00 Chromium	54.94 Manganese	55.85 Iron	58.93 Cobalt	58.69 Nickel	63.55 Copper	65.38 ^{Zinc}	69.72 Gallium	72.64 Germanium	74.92 Arsenic	78.96 Selenium	79.90 Bromine	83.80 Krypton
37 Rb	38 Sr	39 Y	40 Zr	41 dN	42 Mo	43 Tc	44 Ru	45 Rh	46 Pd	47 Ag	Cd 48	49 In	Sn Sn	51 Sb	52 Te	53	54 Xe
85.47 Rubidium	87.61 Strontium	88.91 Yttrium	91.22 Zirconium	92.91 Niobium	95.96 Molybdenum	Technetium	101.1 Ruthenium	102.9 Rhodium	106.4 Palladium	107.9 Silver	112.4 Cadmium	114.8 Indium	118.7 Tin	121.8 Antimony	127.6 Tellurium	126.9 Iodine	131.3 Xenon
55	56	57-71	72	73	74	75	<u>76</u>	LL	78	62	80	₩.	22	83	84	85	86
US 132.9	Ба 137.3		ні 178.5	1a 180.9	W 183.9	ке 186.2	US 190.2	Ir 192.2	Pt 195.1	Au 197.0	нg 200.6	204.4	Pb 207.2	B1 209.0	20	At	Kn
Caesium	Barium	Lanthanoids	Hafnium	Tantalum	Tungsten	Rhenium	Osmium	Iridium	Platinum	Gold	Mercury	Thallium	Lead	Bismuth	Polonium	Astatine	Radon
87 Fr	88 Ra	89–103	104 Rf	105 Db	106 Sg	107 Bh	108 Hs	109 Mt	110 Ds	111 Rg	Cn Cn	113 Nh	114 Fl	115 Mc	116 Lv	117 Ts	118 Og
Francium	Radium	Actinoids	Rutherfordium	Dubnium	Seaborgium	Bohrium	Hassium	Meitnerium I	Darmstadtium	Roentgenium	Copernicium	Nihonium	Flerovium	Moscovium	Livermorium	Tennessine	Oganesson
		Lanthanoids	ids														
		57 1 3	58	59 Dr	09 09	61 Dm	62 Sm	63 Fu	64 64	65 Th	99 Dv	67 Ho	68 Fr	69 1 ^m L	02 47	71 1	
		Lanthanum		9 muin	144.2 Neodymium	L III Promethium	150.4 Samarium	152.0 Europium	157.3 Gadolinium	10 158.9 Terbium	162.5 Dysprosium	164.9 Holmium	167.3 Erbium	168.9 Thulium	173.1 Ytterbium	175.0 Lutetium	
		Actinoids	x		-	-	-	-		-			-	-			
		89 68	06 21	91 Do	92 11	93 Ma	94 Dii	95 Am	96 0	97 10	98 7f	99 E	100 E	101 Md	102 No	103 1 *	
		AC	232.0	га 231.0	238.0	dM	Ги	IIIA				ŝ		nıvı	0N1		
		Actinium	Thorium	Protactinium	Uranium	Neptunium	Plutonium	Americium	Curium	Berkelium	Californium	Einsteinium	Fermium	Mendelevium	Nobelium	Lawrencium	
		Standard ato Elements wi	omic weight: ith no report	Standard atomic weights are abridged to four significant figures. Elements with no reported values in the table have no stable nuclide:	d to four sig the table ha	gnificant figu ve no stable	ures. nuclides.										
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Information on elements with atomic numbers 113 and above is sourced from the International Union of Pure and Applied Chemistry Periodic Table of the Elements (November 2016 version). The International Union of Pure and Applied Chemistry Periodic Table of the Elements (February 2010 version) is the principal source of all other data. Some data may have been modified.